



**Cambridge International Examinations**  
Cambridge International General Certificate of Secondary Education

**CHEMISTRY**

**0620/11**

Paper 1 Multiple Choice (Core)

**May/June 2016**

**45 Minutes**

Additional Materials:      Multiple Choice Answer Sheet  
   Soft clean eraser  
   Soft pencil (type B or HB is recommended)

\* 0 0 2 2 1 4 3 8 7 2 \*

**READ THESE INSTRUCTIONS FIRST**

Write in soft pencil.  
Do not use staples, paper clips, glue or correction fluid.  
Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.  
**DO NOT WRITE IN ANY BARCODES.**

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.  
Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

**Read the instructions on the Answer Sheet very carefully.**

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.  
Any rough working should be done in this booklet.  
A copy of the Periodic Table is printed on page 16.  
Electronic calculators may be used.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **16** printed pages.

1 In which changes do the particles move further apart?

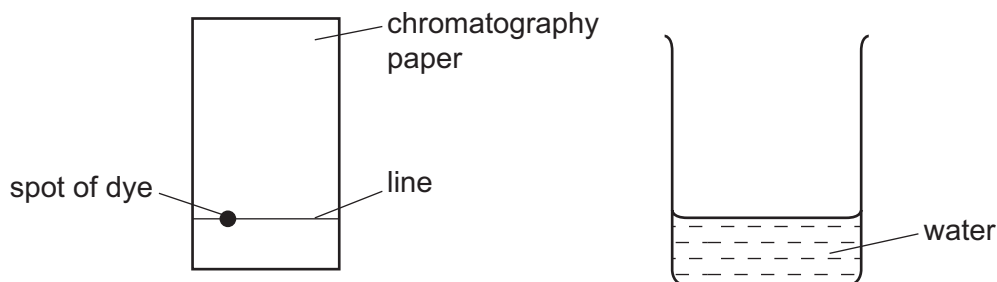


- A** W and X      **B** W and Z      **C** X and Y      **D** Y and Z

2 A sample of a dye is investigated by chromatography.

A line is drawn across a piece of chromatography paper and a spot of the dye is placed on it.

The paper is placed in water.



Which row is correct?

	what is used to draw the line	position of spot
<b>A</b>	ink	above the level of the water
<b>B</b>	ink	below the level of the water
<b>C</b>	pencil	above the level of the water
<b>D</b>	pencil	below the level of the water

3 One of the instructions for an experiment reads as follows.

Quickly add 50 cm<sup>3</sup> of acid.

What is the best piece of apparatus to use?

- A** a burette  
**B** a conical flask  
**C** a measuring cylinder  
**D** a pipette

- 4 Diamond and graphite are macromolecules.

Which statement about diamond and graphite is **not** correct?

- A** They are giant structures with high melting points.  
**B** They are non-conductors of electricity.  
**C** They contain only atoms of a non-metal.  
**D** They have covalent bonds between the atoms.

- 5 The table shows the electronic structure of four atoms.

atom	electronic structure
W	2,8,1
X	2,8,4
Y	2,8,7
Z	2,8,8

Which two atoms combine to form a covalent compound?

- A** W and X      **B** W and Y      **C** X and Y      **D** X and Z

- 6 An aluminium atom has a nucleon number of 27 and a proton number of 13.

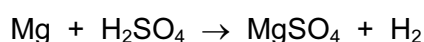
How many neutrons does this aluminium atom contain?

- A** 13      **B** 14      **C** 27      **D** 40

- 7 What happens when a bond is formed between a green gaseous element and a soft metallic element?

- A** The gas atoms gain an electron.  
**B** The gas atoms lose an electron.  
**C** The metal atoms gain an electron.  
**D** The two elements share a pair of electrons.

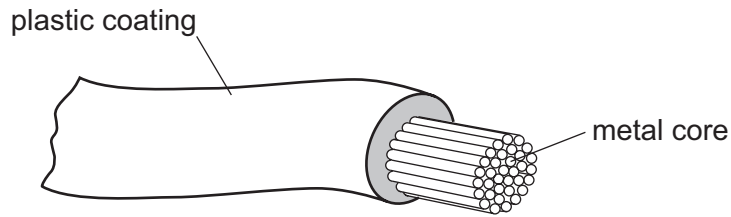
- 8 The equation shows the reaction between magnesium and sulfuric acid.  
 [A<sub>r</sub>: H, 1; O, 16; Mg, 24; S, 32]



In this reaction, which mass of magnesium sulfate is formed when 6 g of magnesium react with excess sulfuric acid?

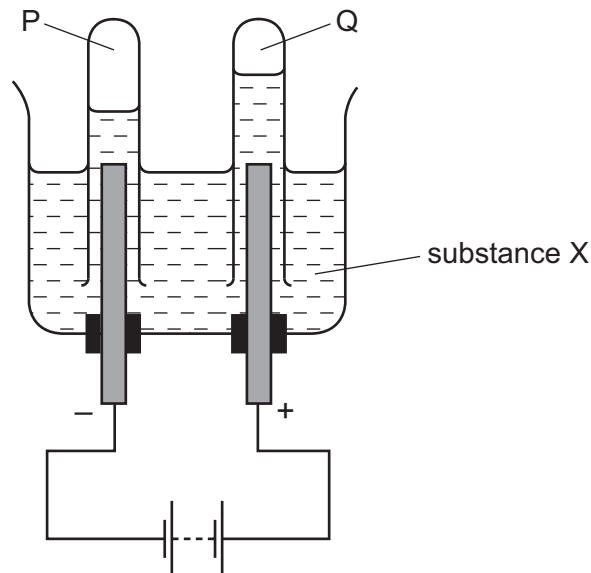
- A** 8      **B** 24      **C** 30      **D** 60

- 9 The diagram shows an electrical cable.



Which statement about the substances used is correct?

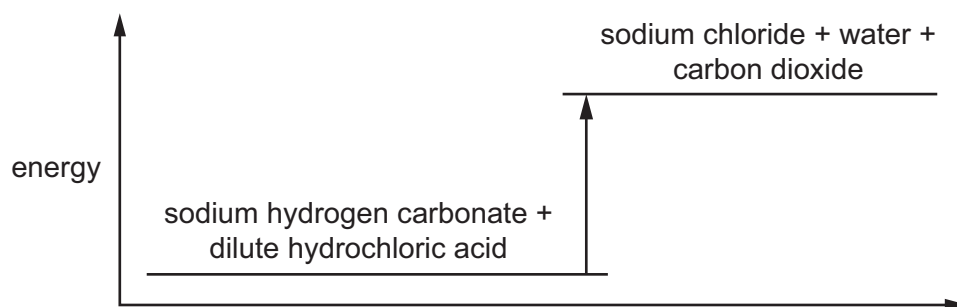
- A** The coating is plastic because it conducts electricity well.  
**B** The core is copper because it conducts electricity well.  
**C** The core is copper because it is cheap and strong.  
**D** The core is iron because it is cheap and strong.
- 10 When substance X is electrolysed, the amount of gases P and Q formed is shown.



What is substance X?

- A** concentrated aqueous sodium chloride  
**B** concentrated hydrochloric acid  
**C** dilute sulfuric acid  
**D** molten lead(II) bromide

- 11 The energy level diagram for the reaction between sodium hydrogen carbonate and dilute hydrochloric acid is shown.

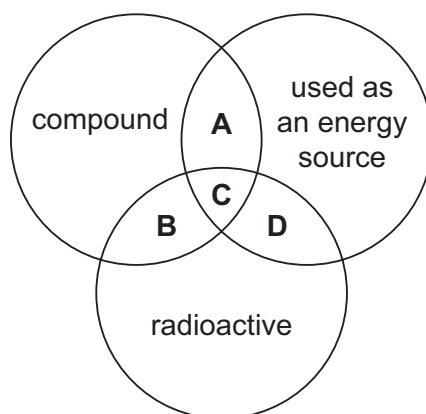


Which row correctly describes the type of reaction and the energy of the reactants and products?

	type of reaction	energy of the reactants and products
<b>A</b>	endothermic	the products have more energy than the reactants
<b>B</b>	endothermic	the reactants have more energy than the products
<b>C</b>	exothermic	the products have more energy than the reactants
<b>D</b>	exothermic	the reactants have more energy than the products

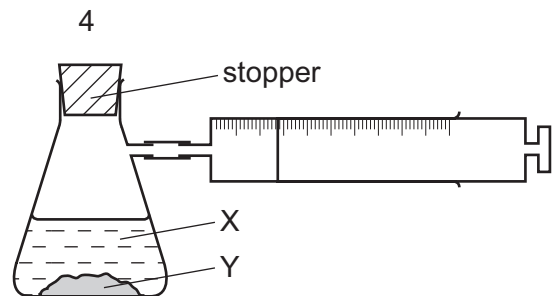
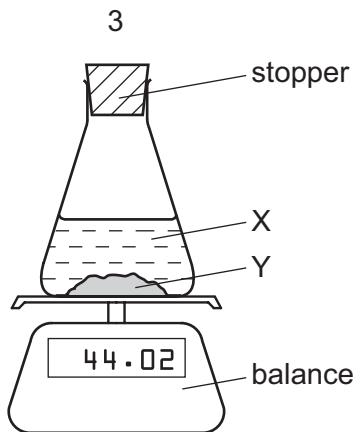
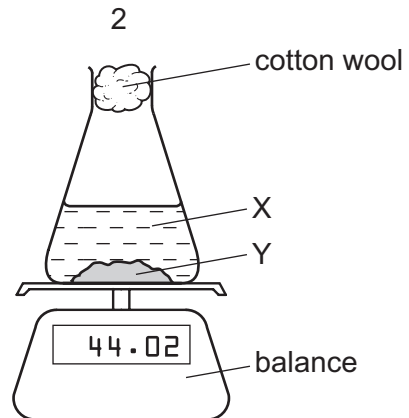
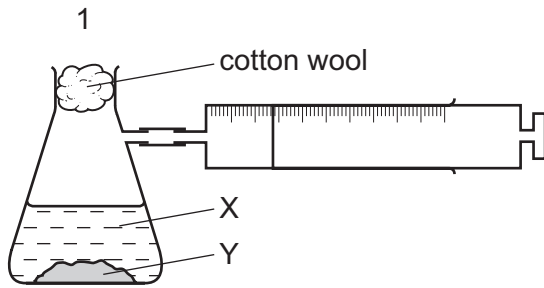
- 12 The diagram shows some properties that substances may have.

To which labelled part of the diagram does  $^{235}\text{U}$  belong?



13 A liquid X reacts with solid Y to form a gas.

Which two diagrams show suitable methods for investigating the rate (speed) of the reaction?



A 1 and 3

B 1 and 4

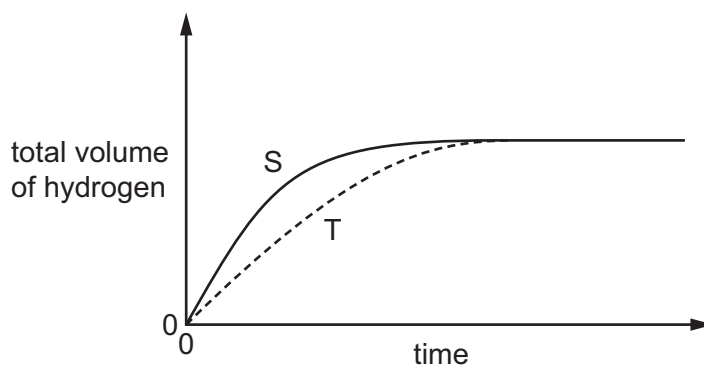
C 2 and 3

D 2 and 4

- 14 An experiment, S, is carried out to measure the volume of hydrogen produced when excess dilute sulfuric acid is added to zinc.

A second experiment, T, is carried out using the same mass of zinc but under different conditions.

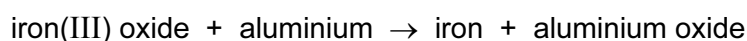
The results of the two experiments are shown.



Which changes in the conditions between experiments S and T give curve T?

	addition of a catalyst	the zinc is in large pieces not powdered
<b>A</b>	✓	✓
<b>B</b>	✓	x
<b>C</b>	x	✓
<b>D</b>	x	x

- 15 Aluminium reacts with iron(III) oxide as shown.



Which statement about this reaction is correct?

- A** Aluminium is oxidised.
- B** Aluminium oxide is reduced.
- C** Iron(III) oxide is oxidised.
- D** Iron is oxidised.
- 16 Which reaction is reversible?
- A**  $\text{Cu} + \text{ZnSO}_4 \rightarrow \text{CuSO}_4 + \text{Zn}$
- B**  $\text{CuO} + \text{H}_2\text{SO}_4 \rightarrow \text{CuSO}_4 + \text{H}_2\text{O}$
- C**  $\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$
- D**  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O} \rightarrow \text{CuSO}_4 + 5\text{H}_2\text{O}$

17 Which statements are properties of an acid?

- 1 reacts with ammonium sulfate to form ammonia
- 2 turns red litmus blue

	1	2
<b>A</b>	✓	✓
<b>B</b>	✓	x
<b>C</b>	x	✓
<b>D</b>	x	x

18 Part of the Periodic Table is shown.

Which element forms an acidic oxide?

The diagram shows a partial periodic table with four groups labeled A, B, C, and D. Group A is the first column on the left. Group B is the first column of the middle section. Group C is the second column of the middle section. Group D is the first column on the right. A small empty box is located above the middle section of the table.

19 A method used to make copper(II) sulfate crystals is shown.

- 1 Place dilute sulfuric acid in a beaker.
- 2 Warm the acid.
- 3 Add copper(II) oxide until it is in excess.
- 4 Filter the mixture.
- 5 Evaporate the filtrate until crystals start to form.
- 6 Leave the filtrate to cool.

What are the purposes of step 3 and step 4?

	step 3	step 4
<b>A</b>	to ensure all of the acid has reacted	to obtain solid copper(II) sulfate
<b>B</b>	to ensure all of the acid has reacted	to remove excess copper(II) oxide
<b>C</b>	to speed up the reaction	to obtain solid copper(II) sulfate
<b>D</b>	to speed up the reaction	to remove excess copper(II) oxide



20 The results of two tests on solid X are shown.

test	observation
aqueous sodium hydroxide added	green precipitate formed
acidified silver nitrate added	yellow precipitate formed

What is X?

- A copper(II) chloride
- B copper(II) iodide
- C iron(II) chloride
- D iron(II) iodide

21 Where in the Periodic Table is the metallic character of the elements greatest?

	left or right side of a period	at the top or bottom of a group
A	left	bottom
B	left	top
C	right	bottom
D	right	top

22 Some properties of four elements, P, Q, R and S, are shown in the table.

Two of these elements are in Group I of the Periodic Table and two are in Group VII.

element	reaction with water	physical state at room temperature
P	reacts vigorously	solid
Q	does not react with water	solid
R	reacts explosively	solid
S	dissolves giving a coloured solution	liquid

Which statement is correct?

- A P is below R in Group I.
- B Q is above R in Group I.
- C Q is below S in Group VII.
- D R is below S in Group VII.

23 Which of the following could be a transition element?

	melting point in °C	density in g/cm <sup>3</sup>	colour	electrical conductor
<b>A</b>	114	4.9	purple	no
<b>B</b>	659	2.7	grey	yes
<b>C</b>	1677	4.5	grey	yes
<b>D</b>	3727	2.3	black	yes

24 Two statements about argon are given.

- 1 Argon has a full outer shell of electrons.
- 2 Argon is very reactive and is used in lamps.

Which is correct?

- A** Both statements are correct and statement 2 explains statement 1.  
**B** Both statements are correct but statement 2 does not explain statement 1.  
**C** Statement 1 is correct but statement 2 is incorrect.  
**D** Statement 2 is correct but statement 1 is incorrect.

25 Three students, X, Y and Z, were told that solid P reacts with dilute acids and also conducts electricity.

The table shows the students' suggestions about the identity of P.

X	Y	Z
copper	iron	graphite

Which of the students are correct?

- A** X, Y and Z      **B** X only      **C** Y only      **D** Z only

26 W, X and Y are metals, one of which is copper and one of which is iron.

- W has a coloured oxide which can be reduced by carbon.
- X has a black oxide and is also found in nature as a pure metal.
- Y has an oxide which cannot be reduced by carbon.

Which metal is the most reactive and what is the possible identity of W?

	most reactive metal	possible identity of W
<b>A</b>	X	Cu
<b>B</b>	X	Fe
<b>C</b>	Y	Cu
<b>D</b>	Y	Fe

27 Tin is a metal that is less reactive than iron and is extracted from its ore cassiterite,  $\text{SnO}_2$ .

Which statements about tin are correct?

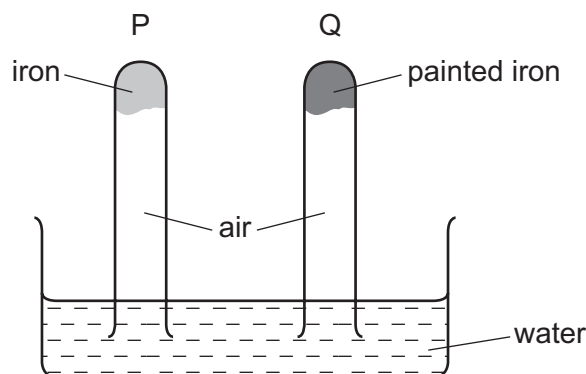
- 1 Tin can be extracted from cassiterite using carbon.
- 2 Tin does not conduct electricity.
- 3 Tin is hard and shiny.

**A** 1, 2 and 3      **B** 1 and 2 only      **C** 1 and 3 only      **D** 2 and 3 only

28 Which statement about the uses of metals is correct?

- A** Aluminium is used in the manufacture of aircraft because of its strength and high density.
- B** Copper is used in electrical wiring because of its strength and high density.
- C** Mild steel is used in the manufacture of car bodies because of its strength and resistance to corrosion.
- D** Stainless steel is used in the construction of chemical plant because of its strength and resistance to corrosion.

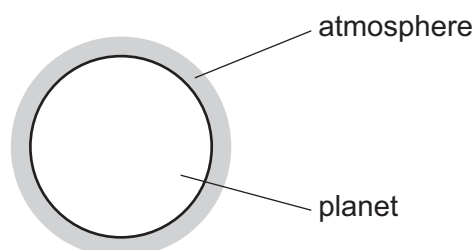
29 The diagram shows an experiment to investigate how paint affects the rusting of iron.



What happens to the water level in tubes P and Q?

	tube P	tube Q
<b>A</b>	falls	rises
<b>B</b>	no change	rises
<b>C</b>	rises	falls
<b>D</b>	rises	no change

30 A new planet has been discovered and its atmosphere has been analysed.



The table shows the composition of its atmosphere.

gas	percentage by volume
carbon dioxide	4
nitrogen	72
oxygen	24

Which gases are present in the atmosphere of the planet in a higher percentage than they are in the Earth's atmosphere?

- A** carbon dioxide and oxygen
- B** carbon dioxide only
- C** nitrogen and oxygen
- D** nitrogen only

- 31 Water was added to separate samples of anhydrous cobalt(II) chloride and anhydrous copper(II) sulfate.

Which row describes the colour changes that take place in these reactions?

	cobalt(II) chloride	copper(II) sulfate
<b>A</b>	blue to pink	blue to white
<b>B</b>	blue to pink	white to blue
<b>C</b>	pink to blue	blue to white
<b>D</b>	pink to blue	white to blue

- 32 Which pollutant found in air does **not** have an effect on respiration?

- A** carbon monoxide
- B** lead compounds
- C** oxides of nitrogen
- D** sulfur dioxide

- 33 A farmer's soil is very low in both nitrogen (N) and phosphorus (P).

Which fertiliser would improve the quality of this soil most effectively?

	percentage		
	nitrogen (N)	phosphorus (P)	potassium (K)
<b>A</b>	11	11	27
<b>B</b>	12	37	10
<b>C</b>	28	10	10
<b>D</b>	31	29	9

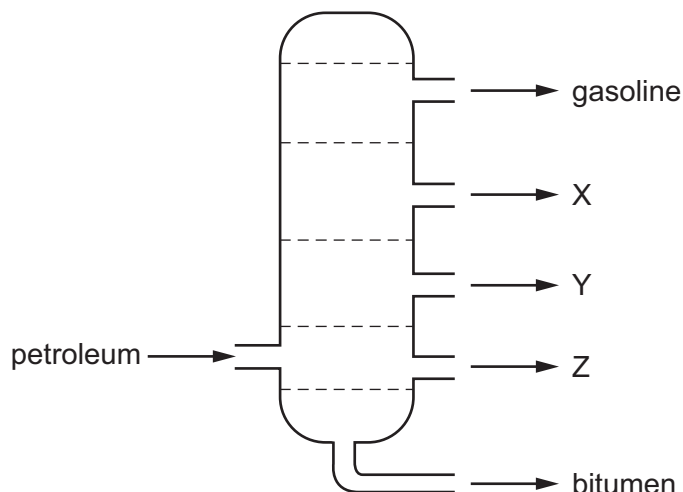
- 34 When limestone is heated it forms lime (calcium oxide) and carbon dioxide.



Which statement is **not** correct?

- A** Carbon dioxide is a greenhouse gas which may contribute to climate change.
- B** Slaked lime is used to neutralise industrial waste.
- C** The lime can be used to treat alkaline soil.
- D** This reaction is an example of thermal decomposition.

35 The diagram shows the separation of petroleum into fractions.



What could X, Y and Z represent?

	X	Y	Z
<b>A</b>	diesel oil	lubricating fraction	paraffin
<b>B</b>	lubricating fraction	diesel oil	paraffin
<b>C</b>	paraffin	lubricating fraction	diesel oil
<b>D</b>	paraffin	diesel oil	lubricating fraction

36 Which of the compounds shown are in the same homologous series?

- 1  $\text{CH}_3\text{OH}$
- 2  $\text{CH}_3\text{CH}_2\text{OH}$
- 3  $\text{CH}_3\text{COOH}$
- 4  $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$

**A** 1, 2 and 3      **B** 1, 2 and 4      **C** 1, 3 and 4      **D** 2, 3 and 4

37 Compound Q decolourises bromine water.

Compound Q has two carbon atoms in each molecule.

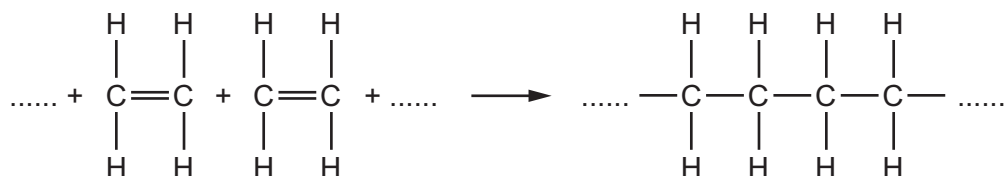
Which statement about compound Q is correct?

- A** It contains carbon-hydrogen double bonds.
- B** It has six hydrogen atoms per molecule.
- C** It has two carbon-carbon double bonds.
- D** It is produced by cracking alkanes.

38 What is used in the production of ethanol from ethene?

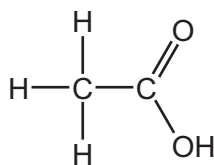
- A hydrogen and oxygen
- B oxygen only
- C steam
- D yeast

39 Ethene forms an addition polymer as shown.



Which terms describe this polymer?

- A a saturated compound called poly(ethane)
  - B a saturated compound called poly(ethene)
  - C an unsaturated compound called poly(ethane)
  - D an unsaturated compound called poly(ethene)
- 40 The diagram shows a molecule of an organic compound W.



Which statement is **not** correct?

- A A solution of W in water has a pH greater than pH 7.
- B A solution of W in water reacts with sodium hydroxide solution.
- C When copper(II) carbonate is added to a solution of W in water, a gas is produced.
- D When magnesium is added to a solution of W in water, a gas is produced.

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The Periodic Table of Elements

		Group																						
I	II	III	IV	V	VI	VII	VIII						VIII											
3 <b>Li</b> lithium 7	4 <b>Be</b> beryllium 9	1 <b>H</b> hydrogen 1	5 <b>B</b> boron 11	6 <b>C</b> carbon 12	7 <b>N</b> nitrogen 14	8 <b>O</b> oxygen 16	9 <b>F</b> fluorine 19	10 <b>Ne</b> neon 20						2 <b>He</b> helium 4										
11 <b>Na</b> sodium 23	12 <b>Mg</b> magnesium 24	<div style="border: 1px solid black; padding: 5px; width: fit-content; margin: auto;"> <p><b>Key</b></p> <p>atomic number</p> <p>atomic symbol</p> <p>name</p> <p>relative atomic mass</p> </div>											13 <b>Al</b> aluminium 27	14 <b>Si</b> silicon 28	15 <b>P</b> phosphorus 31	16 <b>S</b> sulfur 32	17 <b>Cl</b> chlorine 35.5	18 <b>Ar</b> argon 40						36 <b>Kr</b> krypton 84
19 <b>K</b> potassium 39	20 <b>Ca</b> calcium 40	21 <b>Sc</b> scandium 45	22 <b>Ti</b> titanium 48	23 <b>V</b> vanadium 51	24 <b>Cr</b> chromium 52	25 <b>Mn</b> manganese 55	26 <b>Fe</b> iron 56	27 <b>Co</b> cobalt 59	28 <b>Ni</b> nickel 59	29 <b>Cu</b> copper 64	30 <b>Zn</b> zinc 65	31 <b>Ga</b> gallium 70	32 <b>Ge</b> germanium 73	33 <b>As</b> arsenic 75	34 <b>Se</b> selenium 79	35 <b>Br</b> bromine 80	36 <b>Kr</b> krypton 84							
37 <b>Rb</b> rubidium 85	38 <b>Sr</b> strontium 88	39 <b>Y</b> yttrium 89	40 <b>Zr</b> zirconium 91	41 <b>Nb</b> niobium 93	42 <b>Mo</b> molybdenum 96	43 <b>Tc</b> technetium —	44 <b>Ru</b> ruthenium 101	45 <b>Rh</b> rhodium 103	46 <b>Pd</b> palladium 106	47 <b>Ag</b> silver 108	48 <b>Cd</b> cadmium 112	49 <b>In</b> indium 115	50 <b>Sn</b> tin 119	51 <b>Sb</b> antimony 122	52 <b>Te</b> tellurium 128	53 <b>I</b> iodine 127	54 <b>Xe</b> xenon 131							
55 <b>Cs</b> caesium 133	56 <b>Ba</b> barium 137	57–71 lanthanoids	72 <b>Hf</b> hafnium 178	73 <b>Ta</b> tantalum 181	74 <b>W</b> tungsten 184	75 <b>Re</b> rhenium 186	76 <b>Os</b> osmium 190	77 <b>Ir</b> iridium 192	78 <b>Pt</b> platinum 195	79 <b>Au</b> gold 197	80 <b>Hg</b> mercury 201	81 <b>Tl</b> thallium 204	82 <b>Pb</b> lead 207	83 <b>Bi</b> bismuth 209	84 <b>Po</b> polonium —	85 <b>At</b> astatine —	86 <b>Rn</b> radon —							
87 <b>Fr</b> francium —	88 <b>Ra</b> radium —	89–103 actinoids	104 <b>Rf</b> rutherfordium —	105 <b>Db</b> dubnium —	106 <b>Sg</b> seaborgium —	107 <b>Bh</b> bohrium —	108 <b>Hs</b> hassium —	109 <b>Mt</b> meitnerium —	110 <b>Ds</b> darmstadtium —	111 <b>Rg</b> roentgenium —	112 <b>Cn</b> copernicium —	114 <b>Fl</b> flerovium —	116 <b>Lv</b> livermorium —	117 <b>Ts</b> tennessine —	118 <b>Og</b> oganesson —	119 <b>Uue</b> unbinilium —	120 <b>Uub</b> unbinilium —							

57 <b>La</b> lanthanum 139	58 <b>Ce</b> cerium 140	59 <b>Pr</b> praseodymium 141	60 <b>Nd</b> neodymium 144	61 <b>Pm</b> promethium —	62 <b>Sm</b> samarium 150	63 <b>Eu</b> europium 152	64 <b>Gd</b> gadolinium 157	65 <b>Tb</b> terbium 159	66 <b>Dy</b> dysprosium 163	67 <b>Ho</b> holmium 165	68 <b>Er</b> erbium 167	69 <b>Tm</b> thulium 169	70 <b>Yb</b> ytterbium 173	71 <b>Lu</b> lutetium 175
89 <b>Ac</b> actinium —	90 <b>Th</b> thorium 232	91 <b>Pa</b> protactinium 231	92 <b>U</b> uranium 238	93 <b>Np</b> neptunium —	94 <b>Pu</b> plutonium —	95 <b>Am</b> americium —	96 <b>Cm</b> curium —	97 <b>Bk</b> berkelium —	98 <b>Cf</b> californium —	99 <b>Es</b> einsteinium —	100 <b>Fm</b> fermium —	101 <b>Md</b> mendelevium —	102 <b>No</b> nobelium —	103 <b>Lr</b> lawrencium —

lanthanoids

actinoids

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.)